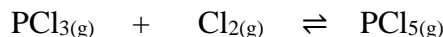


MINISTRY OF EDUCATION, HERITAGE AND ARTS
YEAR 12 CHEMISTRY
REVISION WORKSHEET 5

Write the answers to the following questions in your exercise/activity books.

Strand 3: Reactions

1. Consider the system at equilibrium given below and answer the questions that follow.



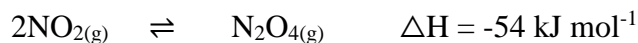
- (i) State how the **concentration of PCl₅ gas** would change if:

- (a) more PCl₃ gas and Cl₂ gas is added. **(1 mark)**
(b) the pressure is decreased. **(1 mark)**

- (ii) Predict the **shift in equilibrium** if:

- (a) a catalyst is added. **(1 mark)**
(b) the pressure is increased. **(1 mark)**

2. Consider the system at equilibrium given below and answer the question that follows.



If the volume of the container is fixed, describe three changes in conditions which would increase the amount of N₂O_{4(g)}. **(3 marks)**

3. Consider the system at equilibrium given below and answer the question that follows.



State how the **concentration of H₂ gas** would change if:

- (i) the temperature is decreased. **(1 mark)**
(ii) CO_(g) is removed from the system as it is formed. **(1 mark)**

THE END